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Multiply molarity by the solution volume (in liters) and the number 106—the molar mass of sodium carbonate—to calculate the mass of sodium carbonate needed. For example, to make 300 mL of 0.2 molar solution, you will need: $0.2 \times 0.3 \text{ L} \times 106 = 6.36 \text{ g}$ Note that $300 \text{ mL} = 0.3 \text{ L}$

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When a solution is made up in a 100 ml volumetric flask, and a 25 ml portion is taken out by pipet, that portion is a one quarter aliquot of the original solution. 2.4 Statistical Aspects of Sampling Uncertainty in environmental measurements can come from both the sampling and the analytical measurement. The total variance is the sum of the two ...

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